



TITLE:

# Preparation of Organo-mercurisulfides. (I)

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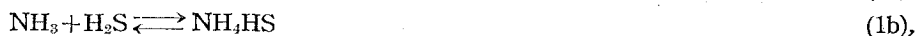
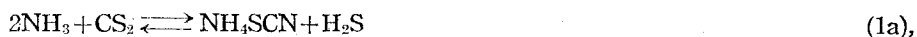
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As the gas in the autoclave is released, the following dissociations take place gradually :



so that when the pressure in the autoclave is released in a hot state, a decrease in the yield of  $\text{NH}_4\text{SCN}$  will be resulted, according to equation (3). Furthermore, the experimental result that by increasing charging density of  $\text{NH}_3$  and  $\text{CS}_2$  the formation of  $\text{NH}_4\text{SCN}$  is decreased can be explained quantitatively with a rate equation derived on above reaction schema from the stand point of chemical kinetics.

## 27. Preparation of Organo-mercurisulfides. (I)

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Merthiolate, *i. e.*, sodium ethylmercurithiosalicylate, which was synthesized by Kharasch *et. al.* in 1926, has strong sterilizing power.

But, the material mercuri-compound of this substance is expensive.

The authors, also with the same purpose, prepared the following compounds, using easily obtainable aromatic mercuri-compounds.

	dec. p.
p-Ethylmercaptomercuribenzoic acid.	179.5°-181°
p-(β-Hydroxyethylmercaptomercuri)-benzoic acid.	168°-170°
p-Carboxylphenylmercurithiosalicylic acid.	215.5°-216.5°
o-(β-Hydroxyethylmercaptomercuri)-benzoic acid.	142°-144°
4-(β-Hydroxyethylmercaptomercuri)-2-chloro-benzoic acid.	147°-149°
3-Chloro-4-carboxylphenylmercuri-thioalicylic acid.	204.5°-205.5°
2-Chloro-tolyl-4-mercurithiosalicylic acid.	153.5°-154.5°
2-Chloro-tolyl-4-mercuripseudothiourea chlorohydrate.	140°-150°